

## Bis(*S*-benzylisothiuronium) tetrachloridozincate(II)

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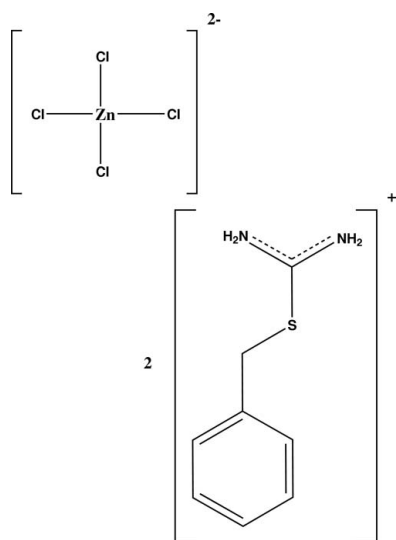
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Key indicators: single-crystal X-ray study;  $T = 293$  K; mean  $\sigma(\text{C}-\text{C}) = 0.004$  Å;  $R$  factor = 0.029;  $wR$  factor = 0.074; data-to-parameter ratio = 22.5.

The asymmetric unit of the title compound,  $(\text{C}_8\text{H}_{11}\text{N}_2\text{S})_2\text{[ZnCl}_4\text{]}$ , contains two *S*-benzylisothiuronium cations which differ in the C—C—S—C torsion angle [165.3 (2) and 81.9 (2)°] and a tetrahedral tetrachloridozincate anion. The crystal structure is stabilized by N—H...Cl, C—H...Cl and C—H...S interactions.

### Related literature

For related literature, see: Hemalatha *et al.* (2006); Zhang *et al.* (1994); Barker & Powell (1998).



### Experimental

#### Crystal data

$(\text{C}_8\text{H}_{11}\text{N}_2\text{S})_2\text{[ZnCl}_4\text{]}$   
 $M_r = 541.67$   
 Monoclinic,  $P2_1/c$   
 $a = 15.2135$  (11) Å  
 $b = 6.4475$  (5) Å  
 $c = 23.9277$  (18) Å  
 $\beta = 95.368$  (1)°

$V = 2336.8$  (3) Å<sup>3</sup>  
 $Z = 4$   
 Mo  $K\alpha$  radiation  
 $\mu = 1.70$  mm<sup>-1</sup>  
 $T = 293$  (2) K  
 $0.27 \times 0.23 \times 0.21$  mm

#### Data collection

Bruker SMART APEX  
 diffractometer  
 Absorption correction: none  
 25129 measured reflections

5481 independent reflections  
 4986 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.026$

#### Refinement

$R[F^2 > 2\sigma(F^2)] = 0.029$   
 $wR(F^2) = 0.073$   
 $S = 1.08$   
 5481 reflections

244 parameters  
 H-atom parameters constrained  
 $\Delta\rho_{\text{max}} = 0.41$  e Å<sup>-3</sup>  
 $\Delta\rho_{\text{min}} = -0.32$  e Å<sup>-3</sup>

**Table 1**

Selected bond lengths (Å).

Cl1—Zn1	2.2792 (5)	Cl3—Zn1	2.2718 (5)
Cl2—Zn1	2.2650 (5)	Cl4—Zn1	2.2589 (5)

**Table 2**

Hydrogen-bond geometry (Å, °).

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
N1—H1A...Cl3	0.86	2.68	3.372 (2)	139
N1—H1B...Cl2 <sup>i</sup>	0.86	2.45	3.255 (2)	157
N2—H2A...Cl4	0.86	2.53	3.219 (2)	138
N2—H2B...Cl3 <sup>ii</sup>	0.86	2.62	3.262 (2)	132
N3—H3A...Cl1 <sup>iii</sup>	0.86	2.72	3.469 (2)	147
N3—H3A...Cl4 <sup>iii</sup>	0.86	2.65	3.244 (2)	128
N3—H3B...Cl1 <sup>iv</sup>	0.86	2.44	3.283 (2)	166
N4—H4A...Cl1 <sup>iii</sup>	0.86	2.49	3.290 (2)	156
N4—H4B...Cl2 <sup>i</sup>	0.86	2.62	3.447 (2)	163
C15—H15A...S1	0.97	2.87	3.591 (2)	132
C15—H15A...Cl2 <sup>v</sup>	0.97	2.77	3.556 (2)	139
C15—H15B...Cl2 <sup>i</sup>	0.97	2.65	3.594 (2)	164

Symmetry codes: (i)  $-x + 1, -y, -z + 1$ ; (ii)  $x, y + 1, z$ ; (iii)  $x - 1, y, z$ ; (iv)  $x - 1, y + 1, z$ ; (v)  $-x + 1, -y + 1, -z + 1$ .

Data collection: SMART (Bruker, 2001); cell refinement: SAINT (Bruker, 2001); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: PLATON (Spek, 2003); software used to prepare material for publication: SHELXL97 and PARST (Nardelli, 1995).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: BT2708).

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**supplementary materials**

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## Bis(*S*-benzylisothiuronium) tetrachloridozincate(II)

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### Comment

Most organic nonlinear optical (NLO) crystals have usually poor mechanical and thermal properties, and are susceptible to damage during processing. It is difficult to grow large optical quality crystals of these materials for device applications (Zhang *et al.*, 1994). For further enhancement of NLO property many efforts have been made on developing new semiorganic NLO materials. The title compound is one of the new metalorganic nonlinear optical crystals. It combines the advantages of both organic and inorganic materials.

The C—N, S—C bond lengths and C—S—C and N—C—N bond angles are comparable with the similar structure reported earlier (Barker & Powell, 1998). There is a difference in the torsion angles C6—C7—S1—C8 [165.3 (2)°] and C14—C15—S2—C16 [81.9 (2)°] in the two molecules which indicates a difference in the conformation of the two molecules in the asymmetric unit.

The crystal structure (Figs. 2 and 3) is stabilized by N—H···Cl, C—H···Cl and C—H···S interactions.

### Experimental

First *S*-benzylisothiuronium chloride (SBTC) was synthesized as discussed in an earlier report (Hemalatha *et al.*, 2006). The solutions of SBTC (5 g m) and zinc chloride (1 g m) were prepared separately in minimum amount of water. Then the solutions were mixed together and stirred for 1 hr at 45°C. The resulting complex was filtered and thoroughly washed with distilled water. The product was recrystallized repeatedly from 0.2 *M* hydrochloric solution to grow transparent and good quality single crystals for NLO applications. Needle shape crystals were obtained from the saturated solution (with water) of the title compound by slow evaporation technique at room temperature.

### Refinement

All H-atoms were refined using a riding model with  $d(\text{C—H}) = 0.93 \text{ \AA}$  or  $d(\text{N—H}) = 0.86 \text{ \AA}$   $U_{\text{iso}} = 1.2U_{\text{eq}}$  (C,N) and  $0.97 \text{ \AA}$ ,  $U_{\text{iso}} = 1.2U_{\text{eq}}$  (C) for CH<sub>2</sub>.

Figures

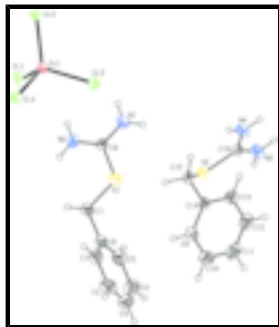


Fig. 1. The molecular structure of title compound, showing 30% probability displacement ellipsoids.

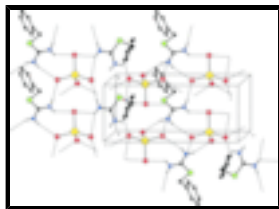


Fig. 2. The molecular packing of the title compound showing N—H...Cl interactions viewed down *c* axis.

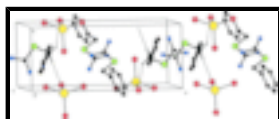


Fig. 3. The molecular packing of the title compound showing C—H...Cl interactions viewed down *c* axis.

**Bis(S-benzylisothiuronium) tetrachloridozincate(II)**

*Crystal data*

(C<sub>8</sub>H<sub>11</sub>N<sub>2</sub>S)<sub>2</sub>[ZnCl<sub>4</sub>]

*M<sub>r</sub>* = 541.67

Monoclinic, *P*2<sub>1</sub>/*c*

Hall symbol: -*P* 2ybc

*a* = 15.2135 (11) Å

*b* = 6.4475 (5) Å

*c* = 23.9277 (18) Å

β = 95.3680 (10)°

*V* = 2336.8 (3) Å<sup>3</sup>

*Z* = 4

*F*<sub>000</sub> = 1104

*D<sub>x</sub>* = 1.540 Mg m<sup>-3</sup>

Mo *K*α radiation

λ = 0.71073 Å

Cell parameters from 2794 reflections

θ = 1.3–25.0°

μ = 1.70 mm<sup>-1</sup>

*T* = 293 (2) K

Needle, colorless

0.27 × 0.23 × 0.21 mm

*Data collection*

Bruker SMART APEX  
diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

*T* = 293(2) K

ω scans

Absorption correction: none

4986 reflections with *I* > 2σ(*I*)

*R*<sub>int</sub> = 0.026

θ<sub>max</sub> = 28.0°

θ<sub>min</sub> = 1.3°

*h* = -20→20

*k* = -8→8

25129 measured reflections  
5481 independent reflections

$l = -30 \rightarrow 31$

### Refinement

Refinement on $F^2$	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.029$	H-atom parameters constrained
$wR(F^2) = 0.073$	$w = 1/[\sigma^2(F_o^2) + (0.0325P)^2 + 1.0433P]$
$S = 1.08$	where $P = (F_o^2 + 2F_c^2)/3$
5481 reflections	$(\Delta/\sigma)_{\max} = 0.001$
244 parameters	$\Delta\rho_{\max} = 0.41 \text{ e } \text{\AA}^{-3}$
Primary atom site location: structure-invariant direct methods	$\Delta\rho_{\min} = -0.32 \text{ e } \text{\AA}^{-3}$
	Extinction correction: none

### Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

### Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

	$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$
S1	0.38843 (3)	0.61059 (12)	0.36896 (3)	0.06677 (19)
N1	0.42614 (13)	0.2947 (3)	0.43326 (8)	0.0562 (5)
H1A	0.4589	0.2133	0.4547	0.067*
H1B	0.3709	0.2676	0.4258	0.067*
C1	0.36865 (17)	1.0458 (4)	0.28381 (11)	0.0598 (6)
H1	0.3774	1.1179	0.3175	0.072*
C2	0.31899 (18)	1.1338 (5)	0.23884 (15)	0.0783 (8)
H2	0.2949	1.2652	0.2423	0.094*
C3	0.3052 (2)	1.0311 (7)	0.19011 (14)	0.0869 (10)
H3	0.2718	1.0920	0.1600	0.104*
C4	0.3395 (2)	0.8398 (7)	0.18442 (11)	0.0867 (10)
H4	0.3293	0.7693	0.1505	0.104*
C5	0.39016 (18)	0.7482 (4)	0.22921 (12)	0.0663 (7)
H5	0.4137	0.6165	0.2252	0.080*
C6	0.40541 (13)	0.8523 (4)	0.27928 (9)	0.0475 (5)
C7	0.46069 (15)	0.7581 (4)	0.32779 (11)	0.0664 (7)

## supplementary materials

H7A	0.4903	0.8661	0.3506	0.080*
H7B	0.5052	0.6680	0.3143	0.080*
C8	0.46009 (12)	0.4585 (3)	0.41203 (8)	0.0399 (4)
S2	0.06186 (3)	0.66804 (7)	0.41032 (2)	0.04038 (11)
N2	0.54328 (12)	0.5038 (3)	0.42243 (9)	0.0555 (5)
H2A	0.5772	0.4246	0.4438	0.067*
H2B	0.5644	0.6130	0.4079	0.067*
C9	0.15114 (15)	0.6634 (4)	0.27656 (10)	0.0578 (6)
H9	0.1742	0.7917	0.2880	0.069*
C10	0.13285 (18)	0.6205 (6)	0.22003 (11)	0.0757 (8)
H10	0.1442	0.7203	0.1936	0.091*
C11	0.09845 (18)	0.4338 (5)	0.20262 (11)	0.0743 (8)
H11	0.0859	0.4075	0.1645	0.089*
C12	0.08244 (17)	0.2855 (5)	0.24096 (10)	0.0651 (7)
H12	0.0588	0.1583	0.2290	0.078*
C13	0.10135 (14)	0.3240 (4)	0.29755 (9)	0.0505 (5)
H13	0.0914	0.2213	0.3235	0.061*
C14	0.13496 (11)	0.5144 (3)	0.31602 (8)	0.0395 (4)
C15	0.15495 (11)	0.5601 (3)	0.37715 (8)	0.0377 (4)
H15A	0.2038	0.6572	0.3818	0.045*
H15B	0.1737	0.4330	0.3964	0.045*
C16	-0.00143 (12)	0.4549 (3)	0.42524 (8)	0.0373 (4)
N3	-0.08381 (12)	0.4941 (3)	0.43308 (8)	0.0527 (4)
H3A	-0.1180	0.3953	0.4418	0.063*
H3B	-0.1038	0.6187	0.4295	0.063*
N4	0.03000 (12)	0.2657 (3)	0.43054 (8)	0.0519 (4)
H4A	-0.0036	0.1658	0.4392	0.062*
H4B	0.0842	0.2415	0.4253	0.062*
C11	0.85432 (4)	-0.02098 (8)	0.44408 (3)	0.05782 (15)
C12	0.76342 (3)	-0.05687 (7)	0.57924 (2)	0.04173 (11)
C13	0.60408 (3)	-0.01885 (8)	0.44970 (3)	0.05317 (14)
C14	0.73327 (3)	0.42677 (7)	0.49034 (2)	0.04099 (11)
Zn1	0.738101 (13)	0.07662 (3)	0.491654 (9)	0.03552 (7)

### Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
S1	0.0303 (2)	0.0860 (4)	0.0846 (4)	0.0098 (3)	0.0085 (2)	0.0446 (4)
N1	0.0452 (10)	0.0590 (11)	0.0640 (12)	-0.0027 (8)	0.0029 (8)	0.0212 (9)
C1	0.0585 (14)	0.0609 (14)	0.0607 (14)	-0.0017 (11)	0.0088 (11)	0.0033 (11)
C2	0.0582 (15)	0.0761 (18)	0.101 (2)	0.0120 (14)	0.0101 (15)	0.0311 (18)
C3	0.0568 (16)	0.128 (3)	0.073 (2)	-0.0069 (18)	-0.0079 (14)	0.039 (2)
C4	0.079 (2)	0.135 (3)	0.0459 (14)	-0.027 (2)	0.0053 (13)	-0.0046 (17)
C5	0.0617 (15)	0.0708 (16)	0.0684 (16)	-0.0033 (12)	0.0167 (12)	-0.0063 (13)
C6	0.0359 (9)	0.0567 (12)	0.0498 (11)	-0.0044 (9)	0.0043 (8)	0.0124 (10)
C7	0.0423 (12)	0.0819 (18)	0.0735 (16)	-0.0119 (11)	-0.0031 (11)	0.0342 (13)
C8	0.0349 (9)	0.0424 (10)	0.0426 (10)	0.0044 (7)	0.0041 (7)	0.0031 (8)
S2	0.0369 (2)	0.0350 (2)	0.0500 (3)	-0.00065 (18)	0.00816 (19)	-0.00128 (19)

N2	0.0386 (9)	0.0523 (10)	0.0724 (13)	-0.0025 (8)	-0.0125 (8)	0.0143 (9)
C9	0.0515 (12)	0.0654 (14)	0.0565 (13)	-0.0130 (11)	0.0048 (10)	0.0116 (11)
C10	0.0635 (15)	0.113 (2)	0.0521 (14)	-0.0118 (16)	0.0110 (12)	0.0252 (15)
C11	0.0585 (15)	0.123 (3)	0.0408 (12)	-0.0074 (16)	0.0039 (11)	-0.0090 (14)
C12	0.0599 (14)	0.0816 (17)	0.0530 (13)	-0.0095 (13)	0.0019 (11)	-0.0212 (12)
C13	0.0502 (12)	0.0544 (12)	0.0466 (11)	-0.0072 (10)	0.0025 (9)	-0.0054 (9)
C14	0.0277 (8)	0.0497 (10)	0.0409 (10)	-0.0010 (7)	0.0021 (7)	-0.0001 (8)
C15	0.0285 (8)	0.0400 (9)	0.0442 (10)	-0.0021 (7)	0.0004 (7)	-0.0029 (7)
C16	0.0372 (9)	0.0397 (9)	0.0351 (9)	-0.0016 (7)	0.0040 (7)	0.0021 (7)
N3	0.0406 (9)	0.0461 (9)	0.0741 (12)	-0.0004 (7)	0.0188 (8)	0.0060 (9)
N4	0.0483 (10)	0.0402 (9)	0.0687 (12)	0.0002 (7)	0.0130 (9)	0.0119 (8)
C11	0.0586 (3)	0.0431 (3)	0.0774 (4)	0.0024 (2)	0.0365 (3)	-0.0026 (2)
C12	0.0459 (2)	0.0367 (2)	0.0417 (2)	-0.00012 (18)	-0.00090 (18)	0.00612 (17)
C13	0.0417 (3)	0.0453 (3)	0.0688 (3)	-0.0069 (2)	-0.0143 (2)	0.0068 (2)
C14	0.0366 (2)	0.0346 (2)	0.0508 (3)	0.00228 (16)	-0.00096 (19)	0.00113 (18)
Zn1	0.03067 (11)	0.03633 (12)	0.03951 (12)	0.00107 (8)	0.00309 (8)	0.00358 (8)

*Geometric parameters (Å, °)*

S1—C8	1.732 (2)	C9—C10	1.383 (4)
S1—C7	1.813 (2)	C9—C14	1.385 (3)
N1—C8	1.300 (3)	C9—H9	0.9300
N1—H1A	0.8600	C10—C11	1.362 (4)
N1—H1B	0.8600	C10—H10	0.9300
C1—C6	1.376 (3)	C11—C12	1.363 (4)
C1—C2	1.378 (4)	C11—H11	0.9300
C1—H1	0.9300	C12—C13	1.380 (3)
C2—C3	1.340 (5)	C12—H12	0.9300
C2—H2	0.9300	C13—C14	1.386 (3)
C3—C4	1.351 (5)	C13—H13	0.9300
C3—H3	0.9300	C14—C15	1.495 (3)
C4—C5	1.391 (4)	C15—H15A	0.9700
C4—H4	0.9300	C15—H15B	0.9700
C5—C6	1.374 (3)	C16—N3	1.309 (2)
C5—H5	0.9300	C16—N4	1.312 (2)
C6—C7	1.497 (3)	N3—H3A	0.8600
C7—H7A	0.9700	N3—H3B	0.8600
C7—H7B	0.9700	N4—H4A	0.8600
C8—N2	1.300 (3)	N4—H4B	0.8600
S2—C16	1.734 (2)	C11—Zn1	2.2792 (5)
S2—C15	1.825 (2)	C12—Zn1	2.2650 (5)
N2—H2A	0.8600	C13—Zn1	2.2718 (5)
N2—H2B	0.8600	C14—Zn1	2.2589 (5)
C8—S1—C7	103.9 (1)	C14—C9—H9	120.1
C8—N1—H1A	120.0	C11—C10—C9	120.8 (3)
C8—N1—H1B	120.0	C11—C10—H10	119.6
H1A—N1—H1B	120.0	C9—C10—H10	119.6
C6—C1—C2	120.6 (3)	C10—C11—C12	120.1 (2)
C6—C1—H1	119.7	C10—C11—H11	119.9



## supplementary materials

C2—C1—H1	119.7	C12—C11—H11	119.9
C3—C2—C1	120.4 (3)	C11—C12—C13	120.0 (2)
C3—C2—H2	119.8	C11—C12—H12	120.0
C1—C2—H2	119.8	C13—C12—H12	120.0
C2—C3—C4	120.6 (3)	C12—C13—C14	120.6 (2)
C2—C3—H3	119.7	C12—C13—H13	119.7
C4—C3—H3	119.7	C14—C13—H13	119.7
C3—C4—C5	120.1 (3)	C9—C14—C13	118.7 (2)
C3—C4—H4	120.0	C9—C14—C15	119.82 (19)
C5—C4—H4	120.0	C13—C14—C15	121.44 (18)
C6—C5—C4	120.0 (3)	C14—C15—S2	113.93 (12)
C6—C5—H5	120.0	C14—C15—H15A	108.8
C4—C5—H5	120.0	S2—C15—H15A	108.8
C5—C6—C1	118.3 (2)	C14—C15—H15B	108.8
C5—C6—C7	121.0 (2)	S2—C15—H15B	108.8
C1—C6—C7	120.6 (2)	H15A—C15—H15B	107.7
C6—C7—S1	108.00 (15)	N3—C16—N4	120.7 (2)
C6—C7—H7A	110.1	N3—C16—S2	115.7 (2)
S1—C7—H7A	110.1	N4—C16—S2	123.5 (2)
C6—C7—H7B	110.1	C16—N3—H3A	120.0
S1—C7—H7B	110.1	C16—N3—H3B	120.0
H7A—C7—H7B	108.4	H3A—N3—H3B	120.0
N1—C8—N2	121.5 (2)	C16—N4—H4A	120.0
N1—C8—S1	116.2 (2)	C16—N4—H4B	120.0
N2—C8—S1	122.3 (2)	H4A—N4—H4B	120.0
C16—S2—C15	104.82 (9)	C14—Zn1—C12	113.28 (2)
C8—N2—H2A	120.0	C14—Zn1—C13	103.76 (2)
C8—N2—H2B	120.0	C12—Zn1—C13	111.95 (2)
H2A—N2—H2B	120.0	C14—Zn1—C11	107.14 (2)
C10—C9—C14	119.7 (2)	C12—Zn1—C11	106.53 (2)
C10—C9—H9	120.1	C13—Zn1—C11	114.24 (3)
C6—C1—C2—C3	0.5 (4)	C14—C9—C10—C11	0.5 (4)
C1—C2—C3—C4	0.2 (5)	C9—C10—C11—C12	-0.7 (5)
C2—C3—C4—C5	-0.5 (5)	C10—C11—C12—C13	-0.2 (4)
C3—C4—C5—C6	-0.1 (4)	C11—C12—C13—C14	1.3 (4)
C4—C5—C6—C1	0.8 (4)	C10—C9—C14—C13	0.6 (3)
C4—C5—C6—C7	-179.1 (2)	C10—C9—C14—C15	179.8 (2)
C2—C1—C6—C5	-1.0 (4)	C12—C13—C14—C9	-1.5 (3)
C2—C1—C6—C7	178.9 (2)	C12—C13—C14—C15	179.3 (2)
C5—C6—C7—S1	-89.8 (2)	C9—C14—C15—S2	91.8 (2)
C1—C6—C7—S1	90.3 (2)	C13—C14—C15—S2	-89.0 (2)
C8—S1—C7—C6	165.34 (19)	C16—S2—C15—C14	81.90 (15)
C7—S1—C8—N1	-159.52 (19)	C15—S2—C16—N3	-159.78 (16)
C7—S1—C8—N2	20.0 (2)	C15—S2—C16—N4	22.5 (2)

### Hydrogen-bond geometry ( $\text{\AA}$ , $^\circ$ )

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
N1—H1A $\cdots$ C13	0.86	2.68	3.372 (2)	139

N1—H1B…C12 <sup>i</sup>	0.86	2.45	3.255 (2)	157
N2—H2A…C14	0.86	2.53	3.219 (2)	138
N2—H2B…C13 <sup>ii</sup>	0.86	2.62	3.262 (2)	132
N3—H3A…C11 <sup>iii</sup>	0.86	2.72	3.469 (2)	147
N3—H3A…C14 <sup>iii</sup>	0.86	2.65	3.244 (2)	128
N3—H3B…C11 <sup>iv</sup>	0.86	2.44	3.283 (2)	166
N4—H4A…C11 <sup>iii</sup>	0.86	2.49	3.290 (2)	156
N4—H4B…C12 <sup>i</sup>	0.86	2.62	3.447 (2)	163
C15—H15A…S1	0.97	2.87	3.591 (2)	132
C15—H15A…C12 <sup>v</sup>	0.97	2.77	3.556 (2)	139
C15—H15B…C12 <sup>i</sup>	0.97	2.65	3.594 (2)	164

Symmetry codes: (i)  $-x+1, -y, -z+1$ ; (ii)  $x, y+1, z$ ; (iii)  $x-1, y, z$ ; (iv)  $x-1, y+1, z$ ; (v)  $-x+1, -y+1, -z+1$ .

Fig. 1

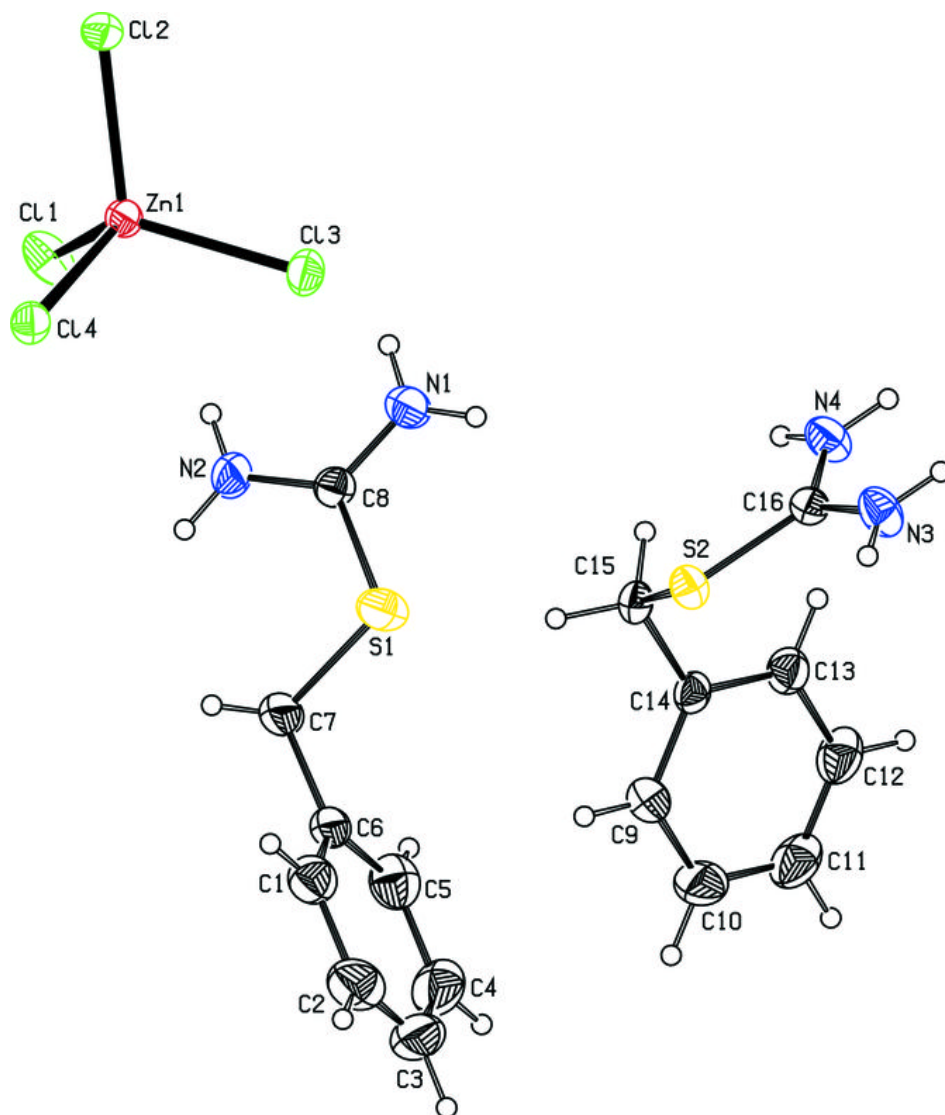


Fig. 2

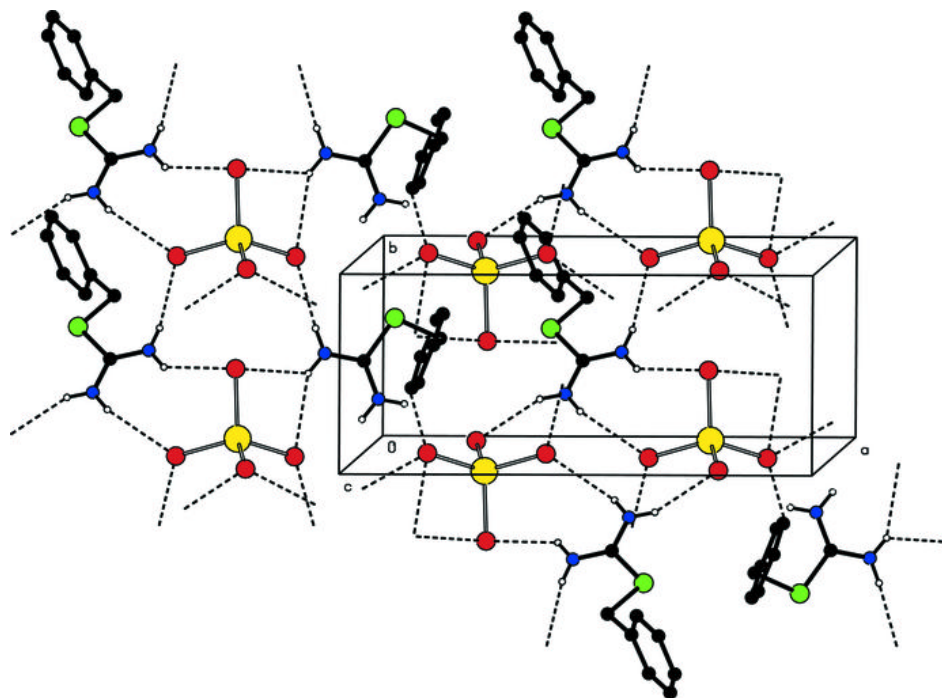


Fig. 3

